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Question: 1.11 - Name the parameters

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that characterise a unit cell. Answer:
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3.1. How would you determine the standard electrode potential of the system $\text{Mg}^{2+} + \text{Mg}$? Ans: A cell will be set up consisting of Mg/MgSO_4 (1 M) as one electrode and standard hydrogen electrode Pt, H_2 (1 atm) $|\text{H}^+$ (1 M) as second electrode, measure the EMF of the cell and also note the direction of deflection in the voltmeter.

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secondary and tertiary alcohols. Ans.
Primary alcohols: (i), (ii), (iii) Secondary
alcohols: (iv), (v)

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Question 14: Suggest two materials
other than hydrogen that can be used as
fuels in fuel cells. Answer: Methane and

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Methanol. Question 15: Explain how rusting of iron is envisaged as setting up of an electrochemical cell. Answer: The chemistry of corrosion is quite complex but it can be considered as an electrochemical phenomenon. Oxidation ...

Class 12 Chemistry Electrochemistry

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